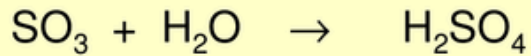


Reactivos	→	Producto
$2Ca + O_2$	→	$2CaO$
$4Al + 3O_2$	→	$2Al_2O_3$
$2Fe + O_2$	→	$2FeO$
$2Co + 3O_2$	→	$2Co_2O_3$



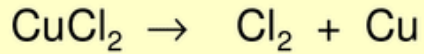
Tipos de reacciones químicas

Reacciones de síntesis



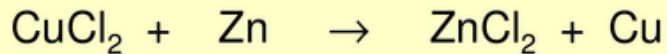
Varias sustancias (elementos o compuestos) se combinan formando una sustancia más compleja.

Reacciones de descomposición



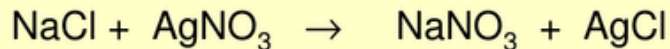
Una sustancia compleja se transforma en otras (elementos o compuestos) de estructura más simple.

Reacciones de sustitución o desplazamiento



Un elemento pasa a ocupar el lugar de otro en un compuesto.

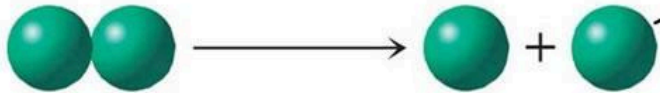
Reacciones de intercambio



Dos elementos de compuestos distintos intercambian sus posiciones obteniendo nuevos compuestos.



combinación



descomposición

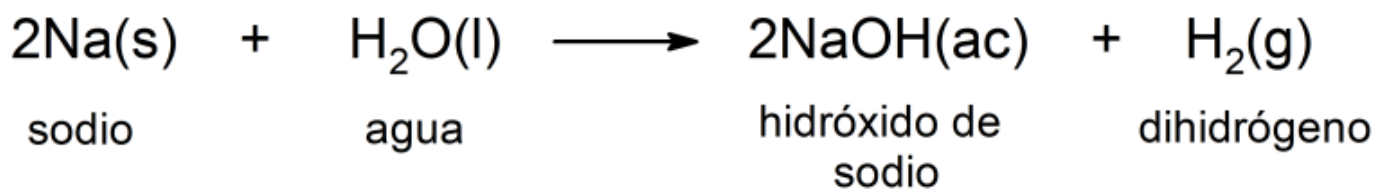
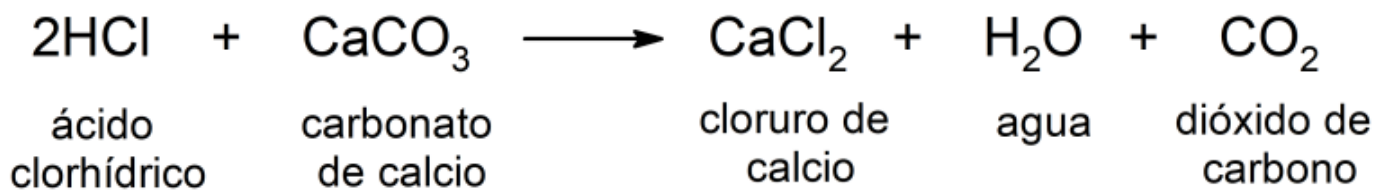
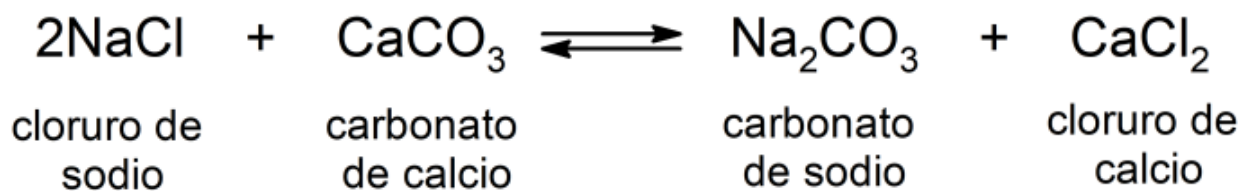
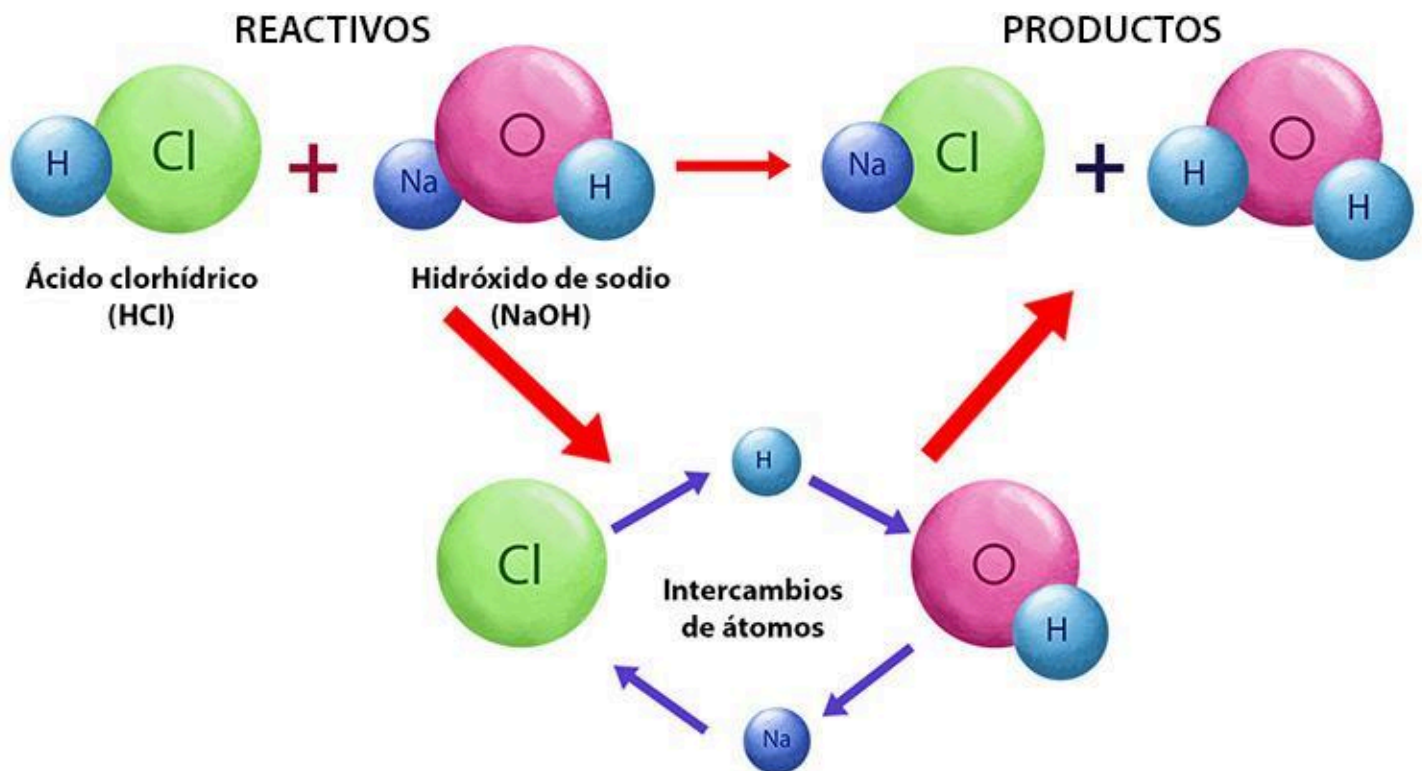


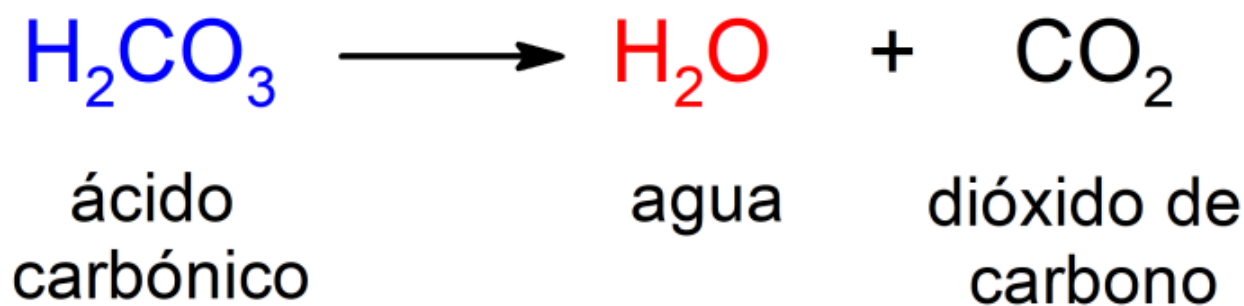
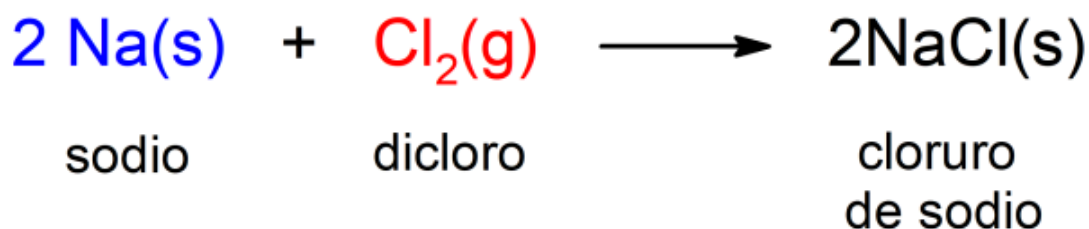
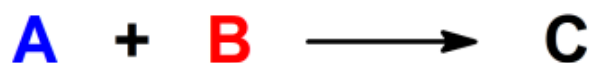
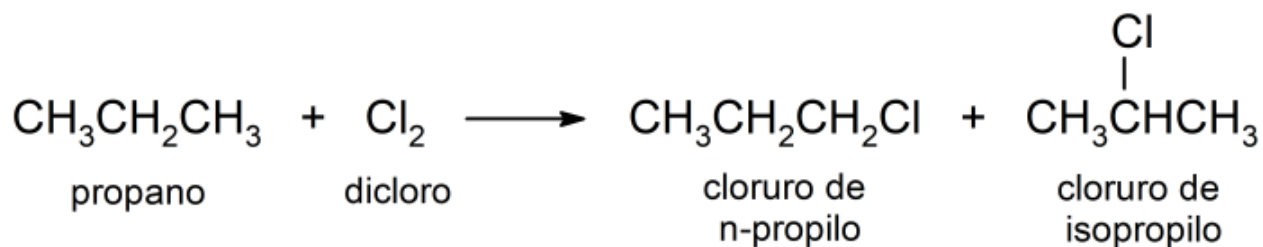
desplazamiento

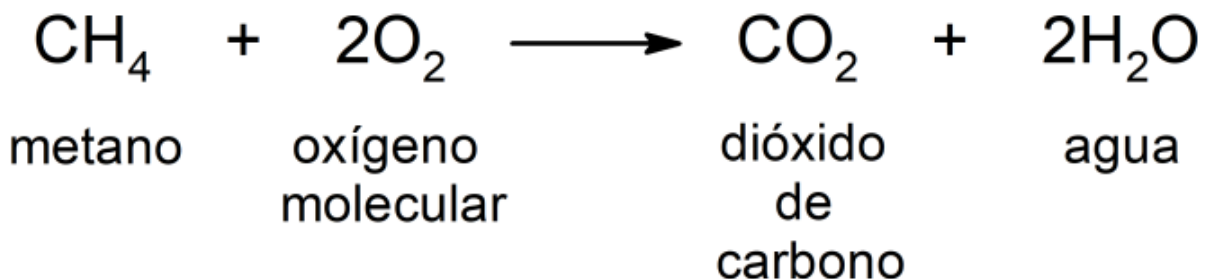
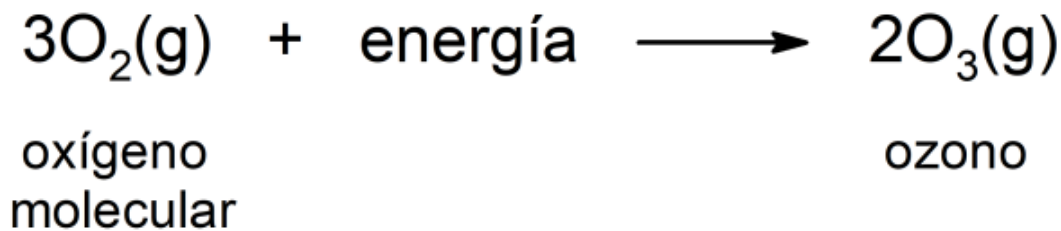
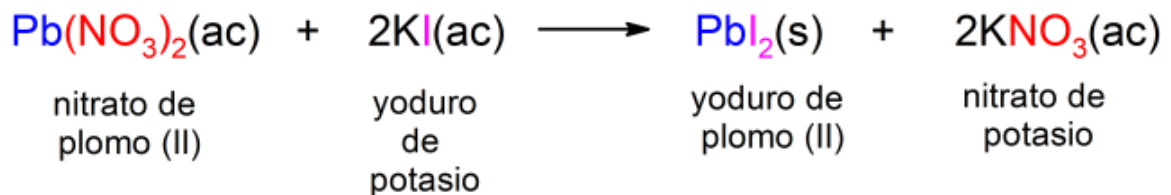
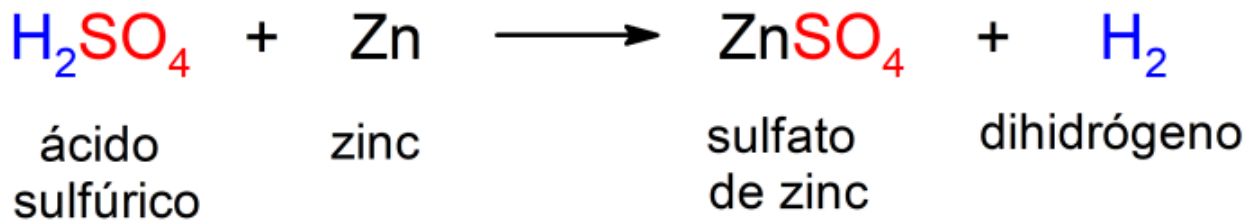
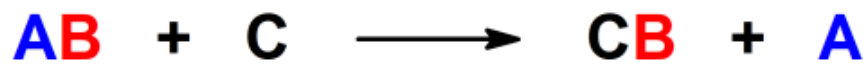


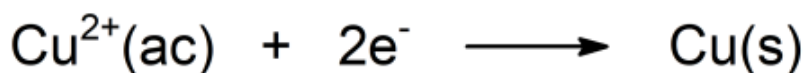
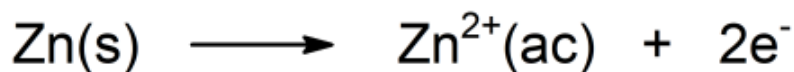
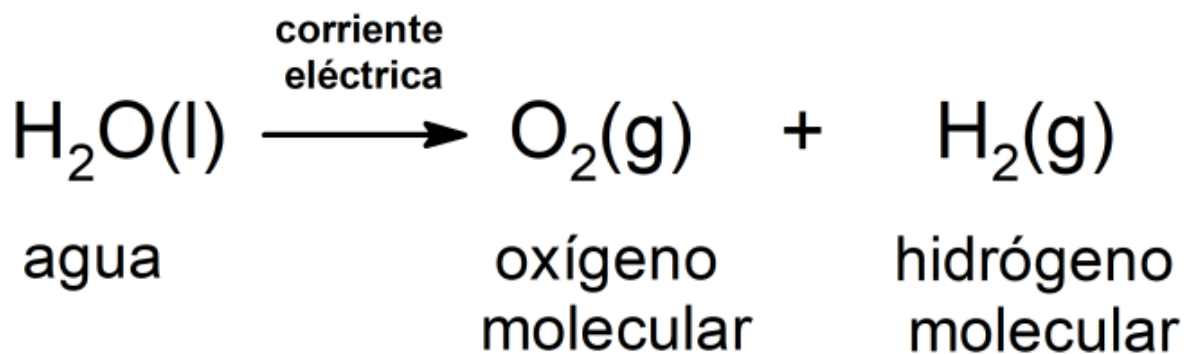
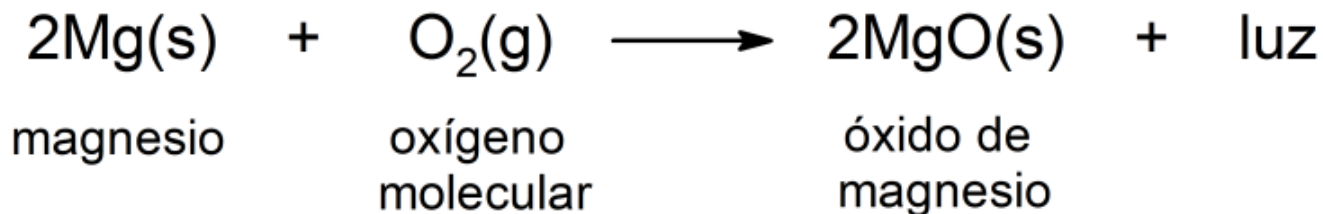
doble desplazamiento

Las esferas representan átomos o grupos de átomos

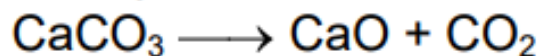
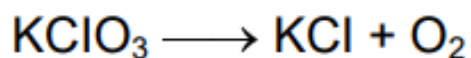




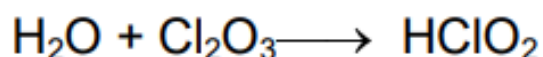




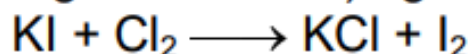
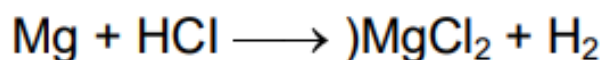
A) REACIONES DE ANÁLISIS O DESCOMPOSICION



B) REACIONES DE SINTESIS O FORMACION



C) REACIONES DE SUSTITUCION SIMPLE



D) REACIONES DE SUSTITUCION DOBLE O METATESIS

